

AP Chemistry Midterm Study Packet

Key Terms:

Chapter 14 – Liquids and Solids

- amorphous solid** — solid in which the arrangement of the representative particles lacks a regular, repeating pattern; do not have a definite melting point
- boiling point** — temperature at which the vapor pressure of a liquid becomes equal to atmospheric pressure
- condensation** — change of state from liquid to gas
- condensed state** — liquid and solid states of matter; states in which a substance has a substantially higher density than in the gaseous state
- covalent network solid** — solid in which strong covalent bonds form a network extending throughout the solid
- crystalline solid** — solid in which the representative particles are in a highly ordered, repeating pattern called a crystal
- deposition** — conversion of a gas directly into a solid without first becoming a liquid
- dipole-dipole force** — intermolecular force of attraction between neighboring permanent dipoles that is created when the dipoles line up so that the positive and negative ends are close to each other
- dispersion force** — intermolecular force of attraction between induced dipoles
- equilibrium vapor pressure** — pressure exerted by a vapor in equilibrium with its liquid; point at which the number of molecules in the vapor state remains solid
- evaporation** — process by which molecules of a liquid escape from the surface of the liquid and enter the gaseous, or vapor, state
- freezing point** — temperature at which the solid and liquid form of a substance exist in equilibrium
- heat of fusion** — heat necessary to convert a given amount of a solid into a liquid
- heat of vaporization** — amount of heat necessary to vaporize a given amount of liquid
- heating curve** — plot of the temperature of a sample as a function of time
- hydrogen bond** — strong intermolecular force between the hydrogen atom of one molecule and a highly electronegative atom of another molecule
- induced dipole** — dipole created by the presence of a neighboring dipole
- intermolecular force** — relatively weak force of attraction that exists between neighboring molecules
- intramolecular force** — force of attraction that exists within a molecule to hold it together
- phase diagram** — plot of the state of a sample as a function of temperature and pressure
- sublimation** — conversion of a solid directly into a gas, without first becoming a liquid
- surface tension** — imbalance of attractive forces at the surface of a liquid that causes the surface to behave as if it had a thin film across it; resistance of a liquid to an increase in the surface area
- vaporization** — change of state from liquid to gas
- viscosity** — resistance to motion that exists between the molecules of a liquid when they move past each other; affected by the extent of the hydrogen bonding

Chapter 15 – Solutions

- alloy** — solid solution in which the atoms of two or more metals are uniformly mixed
- aqueous solution** — solution in which the solvent is water
- boiling point elevation** — colligative property in which the boiling point of a solvent is raised when a nonvolatile solute is dissolved in the solvent; directly related to the concentration of the solution
- colligative property** — property of a solution that depends upon the concentration of the solute, but not its identity
- concentration** — amount of solute dissolved in a given amount of solvent
- freezing point depression** — colligative property in which the freezing point of a solvent is lowered when a nonvolatile solute is dissolved in the solvent; freezing point depression is directly related to the concentration of a solution
- hydration** — process in which water molecules pull solute particles into solution and form a sphere around them
- immiscible** — inability of a liquid to form a solution with another liquid in all proportions
- insoluble** — inability of a substance to dissolve in another substance
- miscible** — ability of a liquid to form a solution with another liquid in all proportions
- molarity** — concentration of a solution determined by the number of moles of solute per kilogram of solvent
- molarity** — concentration of a solution determined by the number of moles of solute per liter of solution
- mass fraction** — concentration of a solution determined by the number of moles of a given solute divided by the total number of moles of solution
- osmosis** — flow of solvent particles from a dilute solution to a concentrated solution
- osmotic pressure** — pressure required to prevent osmosis; arises when the flow of solvent particles from a dilute solution to a concentrated solution across a semipermeable membrane results in uneven heights of the solutions of either side of the membrane
- saturated** — description of a solution that contains as much dissolved solute as it can under existing conditions
- solubility** — amount of solute that dissolves in a given amount of solvent at a given temperature to form a saturated solution
- soluble** — ability of a solution to dissolve in another substance
- solute** — substance that is dissolved in a solvent to form a solution
- solution** — homogeneous mixture of two or more substances in a single physical state
- solvation** — process in which solvent particles pull solute particles into solution and form a sphere around them; called hydration if the solvent is water
- solvent** — substance that does the dissolving in a solution

Chemistry Midterm Study Guide

E. C. Blossey



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