

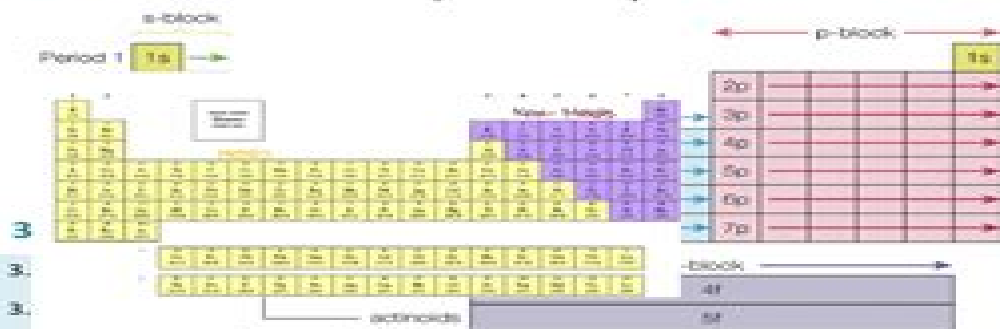
Topic 3: Periodicity

3.1 Periodic Table

- 3.1.1 The periodic table is arranged into four blocks associated with the four sub-levels – s, p, d, and f
- 3.1.2 The periodic table consists of groups (vertical columns) and periods (horizontal rows)
- 3.1.3 The period number (n) is the outer energy level that is occupied by electrons
- 3.1.4 The number of the principal energy level and the number of the valence electrons in an atom can be deduced from its position on the periodic table
- 3.1.5 The periodic table shows the position of metals, non-metals and metalloids
- 3.1.6 Deduction of the electron configuration of an atom from the element's position on the periodic table, and vice versa

Periodic Table Arrangement

- The periodic table is arranged into 4 blocks associated with their last orbital: s, p, d and f
- Elements in the periodic table are arranged by increasing atomic number
- **Group:** A vertical column of elements
 - Group number is the same as the number of valence electrons
- **Period:** A horizontal row of elements
 - Period number is the same number of shells in the atom
 - All shells except the outer shell will be full
- There is a division between metals and non-metals. Metals are on the left and non-metals are on the right
- The farther sub-level occupied is found by the location of the element (which block it is in):



- The location of the following families of the periodic table must be known: alkali metals, alkali earth metals, halogens, noble gases, transition metals, lanthanides and actinoids

- 3.2.1 Atomic radius, ionic radius
- 3.2.2 Electronegativity, ionization energy
- 3.2.3 Oxides change from basic through amphoteric to acidic across a period
- 3.2.4 Prediction and explanation of the metallic and non-metallic behavior of an element based on its position in the periodic table
- 3.2.5

Atomic Radius – The distance from the nucleus to the outermost electron

Ionic Radius – The distance from the nucleus to the outermost electrons in an ion

First ionization energy – The energy required to remove one mole of electrons from one mole of gaseous atoms. This is shown by:

Electron affinity – The energy released when one mole of an electron is added to one mole gaseous atoms. This is shown by:

Electronegativity – A measure of the attraction an atom has for a shared pair of electrons in a covalent bond

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