

Chemistry – Unit 5 Worksheet 2

1. An old (pre-1987) penny is nearly pure copper. If such a penny has a mass of 3.3 g, how many moles of copper atoms would be in one penny?

$$3.3 \text{ g} \times \frac{\# \text{ of moles}}{\# \text{ of grams}} =$$

$$3.3 \text{ g} \times \frac{1 \text{ mol}}{63.55 \text{ g}} = \boxed{0.052 \text{ mol}}$$

2. Four nails have a total mass of 4.42 grams. How many moles of iron atoms do they contain?

$$4.42 \text{ g} \times \frac{\# \text{ of moles}}{\# \text{ of g}} = 4.42 \text{ g} \times \frac{1 \text{ mol}}{55.85 \text{ g/mol}}$$

$$\boxed{0.079 \text{ mol Fe}}$$

3. A raindrop has a mass of 0.050 g. How many moles of water does a raindrop contain?

$$\text{m.m. of water} \quad 2 \times 1.01 \text{ g/mol} + 1 \times 16 \text{ g/mol} = 18.02 \text{ g/mol}$$

$$0.05 \text{ g} \times \frac{1 \text{ mol}}{18.02 \text{ g}} = 0.0028 \text{ mol H}_2\text{O}$$

4. What mass of water would you need to have 15.0 moles of H₂O?

$$\text{m.m. water} = 18.02 \text{ g/mol}$$

$$15 \text{ mol} \times \frac{\# \text{ of g}}{1 \text{ mol}} = 15 \text{ mol} \times \frac{18.02 \text{ g}}{1 \text{ mol}} = \boxed{270.3 \text{ g}}$$

5. One box of Morton's Salt contains 737 grams. How many moles of sodium chloride (NaCl) is this?

m.m. NaCl:

$$\begin{array}{r} \text{Na: } 22.99 \\ \text{Cl: } 35.45 \\ \hline 58.44 \text{ g/mol} \end{array}$$

$$737 \text{ g} \times \frac{1 \text{ mol}}{58.44 \text{ g}} = \boxed{12.61 \text{ mol}}$$

6. A chocolate chip cookie recipe calls for 0.050 moles of baking soda (sodium bicarbonate, NaHCO₃). How many grams should the chef mass out?

m.m. NaHCO₃:

$$\begin{array}{r} 22.99 + 1.01 + \\ 12.01 + 48 = 84.01 \\ \hline \text{g/mol} \end{array}$$

$$0.05 \text{ mol} \times \frac{84.01 \text{ g}}{1 \text{ mol}} = \boxed{4.2 \text{ g NaHCO}_3}$$

7. Rust is iron (III) oxide (Fe₂O₃). The owner of a 1959 Cadillac convertible wants to restore it by removing the rust with oxalic acid, but he needs to know how many moles of rust will be involved in the reaction. How many moles of iron (III) oxide are contained in 2.50 kg of rust?

m.m. Fe₂O₃

$$\begin{array}{r} 2 \times 55.85 = 111.7 \\ 3 \times 16 = 48 \\ \hline 159.7 \end{array}$$

$$2.5 \text{ kg} \times \frac{1000 \text{ g}}{1 \text{ kg}} \times \frac{1 \text{ mol}}{159.7 \text{ g}} = \boxed{15.65 \text{ mol Fe}_2\text{O}_3}$$

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