

# CHEMISTRY REVIEW: KEY CHAPTER 9 – CHEMICAL NAMES AND FORMULAS

- In a Stock name such as iron (II) sulfate, the roman numeral tells us (2)
  - How many atoms of Fe are in one formula unit.
  - How many sulfate ions can be attached to the iron atom.
  - The charge on each Fe ion.
  - The total positive charge of the formula unit.
- The result of changing a subscript in a correctly written chemical formula is to (4)
  - Change the number of moles represented in the formula.
  - Change the charges on the other ions in the compound.
  - Change the formula so that it no longer represents the compound.
  - Have no effect on the formula.
- Derive a generalization for when an acid name will end in the suffix -ic or -ous. If ending in -ic, the anion ended in -ate. If ending in -ous, the anion ended in -ite.
- Derive a generalization for when an acid name will begin with the prefix hydro-. If prefix is hydro-, the anion ended in -ite. This is almost always a monatomic ion.
- Fill in the blanks in the table below.

Compound Name	Formula
Acetic acid	$\text{HC}_2\text{H}_3\text{O}_2$
Aluminum sulfate	$\text{Al}_2(\text{SO}_4)_3$
Lead II chloride	$\text{PbCl}_2$
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Hydrochloric acid	$\text{HCl}$
Dinitrogen pentoxide	$\text{N}_2\text{O}_5$
Copper II carbonate	$\text{CuCO}_3$
Sodium sulfate	$\text{Na}_2\text{SO}_4$
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Mercury sulfate	$\text{HgSO}_4$
Nitrous acid	$\text{HNO}_2$
Magnesium perchlorate	$\text{Mg}(\text{ClO}_4)_2$
Iron II nitrate	$\text{Fe}(\text{NO}_3)_2$
Iron III nitrate	$\text{Fe}(\text{NO}_3)_3$
Sulfur trioxide	$\text{SO}_3$
Sulfur dioxide	$\text{SO}_2$
Barium perchlorate	$\text{Ba}(\text{ClO}_4)_2$
Tetraphosphorus trisulfide	$\text{P}_4\text{S}_3$
Phosphoric acid	$\text{H}_3\text{PO}_4$

- Name the following compound and identify its molecular formula (acid)
  - $\text{Ca}(\text{HCO}_3)_2$  calcium bicarbonate - II
  - $\text{Fe}(\text{NO}_3)_3$  iron III nitrate - II
  - $\text{Fe}(\text{ClO}_4)_3$  iron III perchlorate - II
  - $\text{N}_2\text{O}_5$  dinitrogen pentoxide - III

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