

CHAPTER 4 REVIEW

# Arrangement of Electrons in Atoms

## SECTION 2

**SHORT ANSWER** Answer the following questions in the space provided.

1. \_\_\_\_\_ How many quantum numbers are used to describe the properties of electrons in atomic orbitals?  
 (a) 1 (b) 2 (c) 3 (d) 4
2. \_\_\_\_\_ A spherical electron cloud surrounding an atomic nucleus would best represent  
 (a) an *s*-orbital. (b) a combination of two different *p*-orbitals.  
 (c) a *p*-orbital. (d) a combination of an *s* and a *p*-orbital.
3. \_\_\_\_\_ How many electrons can an energy level of  $n = 4$  hold?  
 (a) 32 (b) 24 (c) 8 (d) 6
4. \_\_\_\_\_ How many electrons can an energy level of  $n = 2$  hold?  
 (a) 32 (b) 24 (c) 8 (d) 6
5. \_\_\_\_\_ Compared with an electron for which  $n = 2$ , an electron for which  $n = 4$  has more  
 (a) spin. (b) energy.  
 (c) particle nature. (d) wave nature.
6. \_\_\_\_\_ According to Bohr, which is the point in the figure below where electrons cannot reside?  
 (a) point A (b) point B (c) point C (d) point D



7. \_\_\_\_\_ According to the quantum theory, point D in the above figure represents  
 (a) the fixed position of an electron.  
 (b) the furthest position from the nucleus that an electron can achieve.  
 (c) a position where an electron probably exists.  
 (d) a position where an electron cannot exist.

# Chapter 4 Arrangement Of Electrons In Atoms Test Answers

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