

The combustion of a hydrocarbon compound according to the reaction equation below provides the maximum theoretical combustion enthalpy of the hydrocarbon per unit mass (standard enthalpy).

- a. (10 marks) Is fuel or oxidiser rich or stoichiometric? (4 marks) Is combustion thermodynamically irreversible? (6 marks) Justify your answer. (4 marks) Is combustion thermodynamically reversible? (6 marks) Justify your answer. (4 marks) Is combustion thermodynamically irreversible? (6 marks) Justify your answer. (4 marks) Is combustion thermodynamically irreversible? (6 marks) Justify your answer. (4 marks)

Using the standard enthalpy of formation of the fuel, calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Using the standard enthalpy of formation of the fuel, calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Using the standard enthalpy of formation of the fuel, calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Using the standard enthalpy of formation of the fuel, calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks)

$$\Delta H^{\circ} = \frac{q}{m} = \frac{-36.7 \text{ kJ}}{-0.001 \text{ kg}} = 36.7 \text{ kJ/kg}$$

- b. (10 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks)

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- c. (10 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks)

Using the standard enthalpy of formation of the fuel, calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Using the standard enthalpy of formation of the fuel, calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Using the standard enthalpy of formation of the fuel, calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Using the standard enthalpy of formation of the fuel, calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks)

- d. (10 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks) Calculate the standard enthalpy of combustion of the hydrocarbon compound. (4 marks)

The standard enthalpy of combustion of a hydrocarbon compound is the standard enthalpy of combustion of the hydrocarbon compound. The standard enthalpy of combustion of a hydrocarbon compound is the standard enthalpy of combustion of the hydrocarbon compound. The standard enthalpy of combustion of a hydrocarbon compound is the standard enthalpy of combustion of the hydrocarbon compound. The standard enthalpy of combustion of a hydrocarbon compound is the standard enthalpy of combustion of the hydrocarbon compound.

$$\Delta H^{\circ} = q + p = \Delta H^{\circ} + p = -36.7 + 0.001 = -36.7 \text{ kJ/kg}$$

# Chem 1050 Homework Exam 1 Assignment Solutions

**RS Peters**



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Center, 1984

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