



pH Properties of Buffer Solutions

AP Chemistry Laboratory #19

Introduction

One of the most important applications of acids and bases in chemistry and biology is that of buffers. A buffer solution resists rapid changes in pH when acids and bases are added to it. Every living cell contains natural buffer systems to maintain the constant pH needed for proper cell function. Many consumer products are also buffered to safeguard their activity. What are buffers made of? How do buffers maintain the delicate pH balance needed for life and health?

Concepts

- Buffer
- Weak acid–weak base
- Conjugate acid–base pair
- Dissociation constant
- Ideal buffer

Background

The ability of buffers to resist changes in pH when acid or base is added is a result of their chemical composition. All buffers contain a mixture of a conjugate acid–base pair; either a weak acid (HA) and its conjugate base (A^-), or a weak base (B), and its conjugate acid (BH^+). Weak acids and weak bases both dissociate slightly in water (Reactions 1 and 2).



These reactions are reversible and both the weak acid and its conjugate base or the weak base and its conjugate acid are present in solution.

The equilibrium constant expressions for these dissociation reactions are:

$$K_a = \frac{[\text{H}_3\text{O}^+][\text{A}^-]}{[\text{HA}]} \quad \text{Equation 1}$$

$$K_b = \frac{[\text{OH}^-][\text{BH}^+]}{[\text{B}]} \quad \text{Equation 2}$$

Buffers control pH because the two buffering components, either HA and A^- or B and BH^+ , are able to neutralize both acids and bases added to the solution.



Chemfax Ph Properties Of Buffer Solutions

J. J. Greczek



Chemfax Ph Properties Of Buffer Solutions:

Buffers for pH and Metal Ion Control D. Perrin, 2012-12-06 This book is intended as a practical manual for chemists biologists and others whose work requires the use of pH or metal ion buffers Much information on buffers is scattered throughout the literature and it has been our endeavour to select data and instructions likely to be helpful in the choice of suitable buffer substances and for the preparation of appropriate solutions For details of pH measurement and the preparation of standard acid and alkali solutions the reader is referred to a companion volume A Albert and E P Serjeant s The Determination of Ionization Constants 1971 Although the aims of the book are essentially practical it also deals in some detail with those theoretical aspects considered most helpful to an understanding of buffer applications We have cast our net widely to include pH buffers for particular purposes and for measurements in non aqueous and mixed solvent systems In recent years there has been a significant expansion in the range of available buffers particularly for biological studies largely in consequence of the development of many zwitterionic buffers by Good et al 1966 These are described in Chapter 3

Buffer Solutions Professor Rob Beynon, J Easterby, 2004-03-01 An indispensable guide to buffers and to understanding the principles behind their use Helps the user to avoid common errors in preparing buffers and their solutions A must for researchers in the biological sciences this valuable book takes the time to explain something often taken for granted buffers used in experiments It answers the common questions such as which buffer should I choose What about the temperature effects What about ionic strength Why is the buffer with the biggest temperature variation used in PCR It provides even the most experienced researchers with the means to understand the fundamental principles behind their preparation and use an indispensable guide essential for everyone using buffers Standard Buffer Solutions George G. Manov, 1947 For the purpose of calibrating pH meters such as the hydrogen calomel and the glass calomel type it is necessary to have on hand a number of buffer solutions of certified pH value covering the ranges of temperature and pH over which it is desired to work Of the large number of compounds both organic and inorganic known to the chemist comparatively few however have the requisite properties such as uniformity and reproducibility of composition freedom from deliquescence or efflorescence ease of preparation and purification and good buffer capacity which would make them acceptable as buffer materials In the past thirty years the concept of the term pH underwent a change from that of the simple definition of $\text{pH} = -\log_{10} [\text{H}^+]$ to the more modern definition of $\text{pH} = -\log_{10} a_{\text{H}^+}$ when it was realized that the potential of the hydrogen electrode was a measure of the activity of the hydrogen ion rather than of its concentration Paradoxically however the activity of single ions cannot be determined from experiment without the use of some simplifying assumptions it is therefore impossible to establish a pH scale on a rigorous thermodynamic basis This point will be discussed later The problem of purification of salts is one primarily of analytical chemistry and the materials can be tested to determine their relative purity and reproducibility by any self consistent method A Study of the PH Values of Buffer Solutions, Using the Hydrogen Electrode and Free Diffusion

Liquid Junctions J. J. Greczek,1965 **Isoelectric Point Determination** Merrell Beyer,1936 *Studying the Ph of Strong Acid, Weak Acid, Salt & Buffer Solutions* Marcia L. Gillette,Norma Robinson North,1998-01-01 **A Comparison of Common Methods of Measuring PH Using Three Different Sets of Buffer Solutions** Roy Ernest Sjogren,1941
Colorimetric Determination of PH Without Buffer Solutions P. H.. Symons,1934 **Polyol-induced PH Changes of Phosphate Buffer Solutions and Their Implications on Food Chemical Stability** Silvia Chuy,2002 **Potentiometric Determination of PH in Different Buffer Solutions at Temperatures of 25, 35, 45, and 60°C Using Poly(acrylic) Acid, Poly(vinyl) Chloride, Poly(methyl) Methacrylate Polymerized, and Bare Tungsten Wire Electrodes** Rasaq O. Babajide,1986 *Crystallization-related PH Changes During Freezing of Sodium Phosphate Buffer Solutions* Gerardo Gómez,1995 **Effect of Composition of Standard Buffer Solutions and Salt Bridge Solutions on the Accuracy and Precision of Blood PH Measurements** Theophil A. Begansky,University of Hartford. Department of Biology,1976 **DIN 19266** ,2000 **PH Measurement and Control** Gregory K. McMillan,1994 This book has been revised to reflect developments in pH control It instructs on the use of graphical and algebraic techniques for system analysis and signal characterization Examples show potential pitfalls and ensure successful system design and implementation The text discusses the factors that make pH loops so difficult to control **Buffers for PH and Metal Ion Control** D. Perrin,1979-04-19 **Buffers for PH and Metal Ion Control** D. D. Perrin,B. Dempsey,1979-05-01 *Solubility Behavior of Pharmaceuticals in Aqueous Solutions* Mohit Chandra Gupta,1952 **Reformulation of Gifford's buffer** Arthur Victor Puccini,1956 **Buffer Solutions for Plating Baths** Louis Latour Watson,1934 *Ph and Buffer Theory* Harry Rilbe,1996

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