

Equilibrium Lab Report

Instructions: For this investigative phenomenon, you will observe the changes that result when you add more of, or take away some of, one of the components of the equilibrium (known as a stress) and determine if there are trends in the behavior of equilibrium systems. Record your observation in the lab report below. You will submit your completed report.

Title: 7.06 Equilibrium Lab Report

Objective(s): Check my understanding of important concepts.

Hypothesis: If the amount of $(\text{Co}(\text{H}_2\text{O})_6^{2+})$ increases, the rate of the (CoCl_4^{2-}) reaction will (decrease) to reach a new equilibrium.

Explain the relationship between forward and reverse reactions at equilibrium and predict how changing the amount of a reactant or product (creating a stress) will affect that relationship.

For example (select one from each underlined section)

If the amount of $(\text{Co}(\text{H}_2\text{O})_6^{2+})$ increases, the rate of the (CoCl_4^{2-}) reaction will (decrease) to reach a new equilibrium.

Procedure:

Access the virtual lab and complete the inquiry experiment. List your controlled variables, independent variable, and dependent variable. Explain why these are the variables. Summarize the steps of the experimental procedure.

Materials:

Equilibrium Virtual Lab

Variables:

Remember, controlled variables are factors that remain the same throughout the experiment. An independent (test) variable changes so that the experimenter can see the effect on other variables. The dependent (outcome) variable will change in response to the test variable.

Controlled variables: Test tubes, solution in test tubes at the start

Independent variable: Everything added into the test tubes

Chemistry 12 Equilibrium Lab Report Answers

Thomas Griffiths



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