

Section 12.2

Chemical Calculations

OBJECTIVES:

- Construct "mole ratios" from balanced chemical equations, and apply these ratios in mole-mole stoichiometric calculations.
- Calculate stoichiometric quantities from balanced chemical equations using units of moles, mass, representative particles, and volumes of gases at STP.

Mole to Mole conversions



- each time we use 2 moles of Al_2O_3 we will also make 3 moles of O_2

These are the two possible conversion factors to use in the solution of the problem.

Mole to Mole conversions

- How many moles of O_2 are produced when 3.34 moles of Al_2O_3 decompose?
$$2\text{Al}_2\text{O}_3 \rightarrow 4\text{Al} + 3\text{O}_2$$

If you know the amount of ANY chemical in the reaction, you can find the amount of ALL the other chemicals!

Practice:



- If 3.84 moles of C_2H_2 are burned, how many moles of O_2 are needed?
- How many moles of C_2H_2 are needed to produce 8.95 mole of H_2O ?

Chemistry Chapter 12 Stoichiometry Notes

J Elliott



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