

Answers to Quick Questions

Chapter 1

- 1 Magnesia, lime, soda, potash and barytes are all the names of compounds.
- 2 Any one of H_2 , N_2 , O_2 or a halogen.
- 3 Any one of the noble gases.
- 4 45
- 5 3
- 6 3
- 7 24, 25 and 26
- 8 The ions represented in Figure 1.6 are formed by the loss of one electron from a magnesium atom. But, the mass of one electron is negligible compared to the mass of the whole atom.
- 9 8 : 1 : 1
- 10 a 12, b 2
- 11 Cu
- 12 Because they are relative to the mass of one twelfth of a carbon-12 atom.
- 13 a 2, b 76.2, c 98.1
- 14 a 134.5, b 148.3, c 277.9
- 15 a 0.25, b 0.1, c 0.1
- 16 a 2.0×10^{-23} g, b 63.5 g, c 51.55 g, d 24.0 g
- 17 a 0.77×10^{23} , b 1.20×10^{23} , c 7.22×10^{24}
- 18 a 0.179 g and 0.445 g, b 4.46×10^{-3} mol and 4.63×10^{-3} mol
c 2.68×10^{21} ions and 2.79×10^{21} ions
- 19 a 28, 29 and 30
b $\frac{92}{100} \times 28 + \frac{5}{100} \times 29 + \frac{3}{100} \times 30 = 25.76 + 1.45 + 0.90 = 28.11$
c They have slightly different percentages of the three isotopes.

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Marco Cascella



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