

14

THE BEHAVIOR OF GASES

Chapter Test II

A. Matching

Match a word or phrase in Column I with the correct description in Column II. Write the letter of the correct answer next to the item.

Column I

1. The constant relating unit temperature changes to changes in volume for a fixed amount of gas at constant pressure (the ratio of the change in volume to the change in temperature).
2. The volume of a fixed amount of a gas at constant pressure in the standard state (0°C and 1 atm).
3. The ratio of the mass of a gas to its volume, independent of the actual amount of gas used.
4. An experimental gas constant determined from the ideal gas law by this equation.
5. A constant of nature that has values of about 8.314 J/mol·K.
6. The a parameter of gas constant equations for determining the gas molecular weight of a gas.
7. An experimental gas constant derived from the ideal gas law using the standard molar volume of a gas at STP.
8. $\frac{P_1 V_1}{T_1} = \frac{P_2 V_2}{T_2}$
9. The amount of heat enough to raise 1 g of water 1°C.
10. A unit of energy (Joules).

Column II

- a. standard gas law
- b. ideal gas constant R
- c. density
- d. compressibility
- e. Boyle's law
- f. molar volume
- g. standard molar volume
- h. volume
- i. Joules/mole
- j. Joules/mole of oxygen

B. Multiple Choice

Choose the correct answer and write its letter on the line.

1. Subliming the solvent of a solid sample for mass determination (drying) requires constant pressure and is:
a. an endothermic process.
b. an exothermic process.
c. an endothermic process.
d. an exothermic process.

Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers

K Morrison



Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers:

Chemistry and Our Universe , In the first of two lectures on the properties of gases review the basic equations that describe their behavior Learn the history of Boyle s law Gay Lussac s law Charles s law and Avogadro s law Then use these four expressions to derive the celebrated ideal gas law

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Table of Contents Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers

1. Understanding the eBook Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - The Rise of Digital Reading Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Personalized Recommendations
 - Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers User Reviews and Ratings
 - Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers and Bestseller Lists
5. Accessing Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers Free and Paid eBooks
 - Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers Public Domain eBooks
 - Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers eBook Subscription Services
 - Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers Budget-Friendly Options
6. Navigating Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers Compatibility with Devices
 - Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Highlighting and Note-Taking Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Interactive Elements Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 8. Staying Engaged with Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 9. Balancing eBooks and Physical Books Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Setting Reading Goals Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Fact-Checking eBook Content of Chemistry Chapter 14 The Behavior Of Gases Worksheet Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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