

Unit  
**08**

## Chemical Reactivity

### Long Answer Questions

**Q1. Define metals. Write down physical and chemical properties of metals.**

**Ans. Metals**

Metals are the elements which are electropositive and form cations by losing electrons. Metals can be categorized.

- a. **Very reactive:** potassium, sodium, calcium, magnesium and aluminium
- b. **Moderately reactive:** Zinc, iron, tin and lead.
- c. **Least reactive or noble metals:** copper, mercury, silver and gold.

**Physical Properties of Metals**

- i. Almost all metals are solids except mercury.
- ii. Metals have high melting and boiling points.
- iii. Metals possess metallic luster and can be polished.
- iv. Metals are malleable (can be hammered into sheets), ductile (can be drawn into wires) and give off a tone when hit.
- v. Metals are good conductor of heat and electricity.
- vi. Metals have high density.
- vii. Metals are hard (except sodium and potassium).

**Chemical Properties of Metals**

- i. Metals easily lose electrons and form positive ions.
- ii. Metals readily react with oxygen to form basic oxides.
- iii. Metals usually form ionic compounds with non metals.
- iv. Metals have metallic bonding.

**Q2. Write a note on electropositive character of metals.**

**Ans. Electropositive character or Metallic character**

Metals have the tendency to lose their valence electrons. This property of a metal is termed as electropositivity or metallic character.

**Explanation**

A metal lose its electrons more easily, therefore more electropositive it will be. The number of electrons lost by an atom of a metal is called its valency. For example, sodium atom can lose 1 electron to form a positive ion.

# Chemistry Chapter 8

**Lei Huang**

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