

Problem 3

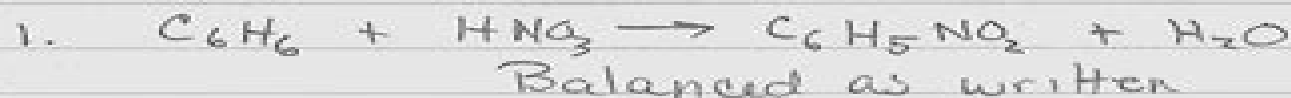


(10 pts)

$$\begin{aligned}
 8 \times C &= 8 \times 12.01 = 96.08 & \% &= \frac{96.08}{194.19} \times 100 = 49.5\% C \\
 10 \times H &= 10 \times 1.00794 = 10.08 & \% &= \frac{10.08}{194.19} \times 100 = 5.2\% H \\
 4 \times N &= 4 \times 14.0067 = 56.03 & \% &= \frac{56.03}{194.19} \times 100 = 28.9\% N \\
 2 \times O &= 2 \times 15.9994 = 31.9988 & \% &= \frac{31.9988}{194.19} \times 100 = 16.5\% O \\
 &&& 194.19
 \end{aligned}$$

Problem 4

5 pts



(10 pt)

$$2. \quad 22.4 \text{ g } C_6H_6 \times \frac{1 \text{ mol } C_6H_6}{78.11 \text{ g } C_6H_6} \times \frac{1 \text{ mol } C_6H_5NO_2}{1 \text{ mol } C_6H_6} \times \frac{123.1052 \text{ g } C_6H_5NO_2}{1 \text{ mol } C_6H_5NO_2}$$

$$\text{Theoretical Yield} \longrightarrow = \underline{\underline{35.28 \text{ g } C_6H_5NO_2}}$$

(10 pts)

$$3. \quad \% \text{ Yield} = \frac{\text{Experimental (Actual)}}{\text{Theoretical}} \times 100$$

$$= \frac{31.6}{35.28} \times 100 = \underline{\underline{89.6\%}}$$

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