

Percent Yield Practice I

1. A student uses stoichiometry to calculate the predicted yield of CO_2 in a chemical reaction as 323 g. When the student *mixes the chemicals* to make the CO_2 , he finds that the reaction only produces *308 g* of CO_2 . What is the percent yield?

$$\text{Percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$$

$$\text{Percent yield} = \frac{308}{323} \times 100$$

$$\text{Percent yield} = 95.4\%$$

Chemistry Theoretical And Percent Yield Answers

**John W. Moore, Conrad L.
Stanitski, Peter C. Jurs**



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